



# Standard Test Methods for Chemical Analysis of Gypsum and Gypsum Products [Metric]<sup>1</sup>

This standard is issued under the fixed designation C 471M; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ε) indicates an editorial change since the last revision or reapproval.

## 1. Scope \*

1.1 These test methods cover the chemical analysis of gypsum and gypsum products, including gypsum ready-mixed plaster, gypsum wood-fibered plaster and gypsum concrete.

NOTE 1—Gypsum reference standard materials, prepared by Domtar, Inc are available through Brammer Standards Company, Inc.

1.2 The test methods appear in the following order:

	Sections
Alternative Procedure for Analysis for Calcium Sulfate by Ammonium Acetate Method	17-22
Alternative Procedure for Analysis for Sodium Chloride by the Coulometric Method	23-29
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1.3 The values stated in SI units are to be regarded as the standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

## 2. Referenced Documents

### 2.1 ASTM Standards:

- C 11 Terminology Relating to Gypsum and Related Gypsum Building Materials and Systems<sup>2</sup>
- C 22/C 22M Specification for Gypsum<sup>2</sup>
- C 28 Specification for Gypsum Plasters<sup>2</sup>
- C 59 Specification for Gypsum Casting and Molding Plaster<sup>2</sup>
- C 61 Specification for Gypsum Keene's Cement<sup>2</sup>
- C 317 Specification for Gypsum Concrete<sup>2</sup>

<sup>1</sup> These test methods are under the jurisdiction of ASTM Committee C-11 on Gypsum and Related Building Materials and Systems and are the direct responsibility of Subcommittee C11.01 on Specifications and Test Methods for Gypsum Products.

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<sup>2</sup> *Annual Book of ASTM Standards*, Vol 04.01.

C 842 Specification for Application of Interior Gypsum Plaster<sup>2</sup>

D 1193 Specification for Reagent Water<sup>3</sup>

D 1428 Methods of Test for Sodium and Potassium in Water and Water-Formed Deposits by Flame Photometry<sup>3</sup>

D 2013 Method of Preparing Coal Samples for Analysis<sup>4</sup>

E 11 Specification for Wire-Cloth Sieves for Testing Purposes<sup>5</sup>

## 3. Terminology

3.1 *Definitions*—Definitions shall be in accordance with Terminology C 11.

3.2 *Definitions of Terms Specific to This Standard:*

3.2.1 *sample as received, n*—a representative portion of raw gypsum or gypsum product in the state received by the testing laboratory, including aggregates, impurities and water content.

3.2.2 *dried sample, n*—a sample dried of free water.

3.2.3 *riffle, n*—a hand feed sample divider device that divides the sample into two parts of approximately the same weight. **(D 2013)**

## 4. Preparation of Sample

4.1 *General Procedures*—Details of sample preparation will vary according to the type of material being tested.

4.1.1 *Sample As Received*—Use a sufficient amount of sample such that, after sieving, not less than 50 g of sample will remain for testing. Weigh the entire sample immediately after opening the container in which the material was received. This will become the weight of the sample as received.

4.1.2 *Drying*—Dry the sample in accordance with Section 7. This will be the weight of the dried sample.

4.1.3 *Crushing and Grinding*—Crush and grind the sample by hand with a mortar and pestle or by mechanical crushing and grinding equipment to pass a 250-µm (No. 60) sieve. Take care, particularly with mechanical equipment, not to expose the sample to temperatures of more than 52°C. Clean the equipment thoroughly between samples. Thoroughly remix the ground sample and store it in an airtight container to avoid contamination.

<sup>3</sup> *Annual Book of ASTM Standards*, Vol 11.01.

<sup>4</sup> *Annual Book of ASTM Standards*, Vol 05.05.

<sup>5</sup> *Annual Book of ASTM Standards*, Vol 14.02.

\*A Summary of Changes section appears at the end of this standard.

4.1.4 *Rehydrating*—Thoroughly blend and rehydrate samples which contain calcium sulfate in forms other than  $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$  and natural anhydrite. Place the sample in distilled water and keep it wet for not less than 48 h. Dry the hydrated sample in an oven at  $45 \pm 3^\circ\text{C}$  to constant weight and recrush or grind it in accordance with 4.1.3.

4.1.5 *Sample Reduction*—Thoroughly mix and reduce large samples as required by quartering or by the use of a riffle to obtain a specimen of approximately 50 g.

4.2 *Gypsum (C 22/C 22M)*—Gypsum samples will be received in the form of rocks, powder or both. If necessary reduce and crush the entire dried sample in accordance with 4.1.3 and 4.1.5.

4.3 *Gypsum Plaster, (C 28)*.

4.3.1 *Gypsum Ready-Mixed Plaster or Gypsum WoodFibered Plaster*—Screen the dried sample through a 150- $\mu\text{m}$  (No. 100) sieve<sup>6</sup> and discard the residue retained on the sieve. Reweigh the remaining sample and calculate the percentage of the dried sample. Reduce the sample in accordance with 4.1.5. Thoroughly blend and rehydrate the specimen in accordance with 4.1.4.

4.3.2 *Gypsum Neat Plaster or Gypsum Gauging Plaster*—Reduce the dried sample in accordance with 4.1.5. Thoroughly blend and rehydrate the specimen in accordance with 4.1.4.

4.4 *Gypsum Casting and Molding Plaster, (C 59)*—Reduce the dried sample in accordance with 4.1.5. Thoroughly blend and rehydrate the specimen in accordance with 4.1.4.

4.5 *Gypsum Keene's Cement, (C 61)*—Reduce the dried sample in accordance with 4.1.5. Blend in no more than 1 % molding plaster or  $\text{K}_2\text{SO}_4$  and rehydrate the specimen in accordance with 4.1.4.

4.6 *Gypsum Concrete, (C 317)*—Screen the dried sample through a 150- $\mu\text{m}$  (No. 100) sieve<sup>6</sup> and discard the residue retained on the sieve. Reweigh the remaining sample and calculate the percentage of the dried sample. Reduce the sample in accordance with 4.1.5. Thoroughly blend and rehydrate the specimen in accordance with 4.1.4.

4.7 *Gypsum Board*—Cut or break the dried sample into small pieces. Using a mortar and pestle, strike the pieces of the sample to loosen the paper face. Remove the pieces of paper by hand as they are separated from the core of the gypsum board. Carefully scrape any remaining powder from the paper. When all the paper has been removed from the pieces of the sample, reduce the sample in accordance with 4.1.5. Thoroughly blend and rehydrate the specimen in accordance with 4.1.4.

## COMPLETE PROCEDURE

### 5. Apparatus

5.1 *Analytical Balance*—Capable of weighing not less than 1 g at a precision of 0.0001 g.

5.2 *Balance*—Capable of weighing not less than 100 g at a precision of 0.001 g.

5.3 *Drying Oven*—A mechanical convection oven set at  $45 \pm 3^\circ\text{C}$ .

5.4 *Desiccator*—Capable of being tightly sealed and containing calcium chloride or equivalent desiccant.

5.5 *Calcining Oven or Furnace*—Capable of achieving and maintaining temperatures to not less than  $1000^\circ\text{C}$ .

5.6 *Weighing Bottles*—Borosilicate glass or ceramic containers with lids that can be sealed tightly.

5.7 *Hot Plate*—A controllable hot plate capable of heating casseroles to approximately  $120^\circ\text{C}$ .

5.8 *Porcelain Casseroles*—With a capacity of 50 to 100 mL.

5.9 *Filtering Funnels*.

5.10 *Filter Paper*—Ashless filter paper Whatman #42 or equivalent.<sup>7</sup>

5.11 *Porcelain Crucibles*.

5.12 *Mortar and Pestle*.

5.13 *Mechanical Jaw Crusher*—Capable of crushing gypsum rocks up to 50 mm diameter.

5.14 *Mechanical Grinder*—Burr mill or equivalent capable of grinding the granular output of the jaw crusher specified in 5.13.

### 6. Reagents

6.1 *Purity of Reagents*—Use reagent grade chemicals in all tests. Unless otherwise indicated, use reagents that conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.<sup>8</sup> If it is necessary to use other grades first ascertain that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.1.1 *Ammonium Chloride* ( $\text{NH}_4\text{Cl}$ ).

6.1.2 *Ammonium Hydroxide* (sp gr 0.90)—Concentrated ammonium hydroxide ( $\text{NH}_4\text{OH}$ ).

6.1.3 *Ammonium Nitrate* (25 g/L)—Dissolve 25 g of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) in water and dilute to 1 L.

6.1.4 *Ammonium Oxalate* ( $(\text{NH}_4)_2\text{C}_2\text{O}_4$ ).

6.1.5 *Barium Chloride* (100 g/L)—Dissolve 100 g of barium chloride ( $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ ) in water and dilute to 1 L.

6.1.6 *Calcium Chloride* ( $\text{CaCl}_2$ )—Anhydrous Calcium Chloride with a combined water of not more than 5 %.

6.1.7 *Hydrochloric Acid* (sp gr 1.19)—Concentrated hydrochloric acid (HCl).

6.1.8 *Hydrochloric Acid* (1 + 4)—Mix 1 volume of HCl (sp gr 1.19) with 4 volumes of water.

6.1.9 *Hydrochloric Acid* (1 + 5)—Mix 1 volume of HCl (sp gr 1.19) with 5 volumes of water.

6.1.10 *Nitric Acid* (sp gr 1.42)—Concentrated nitric acid ( $\text{HNO}_3$ ).

6.1.11 *Potassium Chromate Solution* (100 g/L)—Dissolve 5 g of potassium chromate ( $\text{K}_2\text{CrO}_4$ ) in 50 mL of water, mix, add 10 drops of 0.05 N silver nitrate ( $\text{AgNO}_3$ ) solution, allow to stand for 5 min, and filter.

6.1.12 *Potassium Permanganate* (5.6339 g/L)—Dissolve 5.6339 g of potassium permanganate ( $\text{KMnO}_4$ ) in water and dilute to 1 L.

<sup>7</sup> Whatman No. 42 or an equivalent filter paper has been found suitable for this purpose.

<sup>8</sup> *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopoeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

<sup>6</sup> Detailed requirements for this sieve are given in Specification E 11.

6.1.13 *Silver Nitrate, Standard Solution (0.05 N)*—Prepare and standardize a 0.05 N silver nitrate ( $\text{AgNO}_3$ ) solution.

6.1.14 *Sodium Ammonium Phosphate*—( $\text{NaNH}_4\text{HPO}_4$ ).

6.1.15 *Sulfuric Acid* (sp gr 1.84)—Concentrated sulfuric acid ( $\text{H}_2\text{SO}_4$ ).

6.1.16 *Sulfuric Acid (1 + 6)*—Carefully mix 1 volume of  $\text{H}_2\text{SO}_4$  (sp gr 1.84) with 6 volumes of water.

6.1.17 *Nitric Acid (0.1 N)*—Mix 1.4 mL of  $\text{HNO}_3$  (sp gr 1.42) with 200 mL of water.

6.1.18 *Phenolphthalein Indicator Solution*—Dissolve 0.25 g of phenolphthalein in 30 mL of methanol and dilute to 50 mL with water.

6.1.19 *Sodium Hydroxide Solution (0.1 N)*—Dissolve 1 g of sodium hydroxide ( $\text{NaOH}$ ) in 250 mL of water.

6.1.20 *Water*—Reagent water shall be in accordance with Specification D 1193, type II. Specification D 1193 gives the following values for type II grade water.

Electrical conductivity, max, $\mu\text{S}/\text{cm}$ at 298 K (25-C)	1.0
Electrical resistivity, min, $\text{M}\Omega\cdot\text{cm}$ at 298 K (25-C)	1.0
Total organic carbon (TOC), max, $\mu\text{g}/\text{L}$	50.0
Sodium, max, $\mu\text{g}/\text{L}$	5.0
Chlorides max, $\mu\text{g}/\text{L}$	5.0
Total silica, max, $\mu\text{g}/\text{L}$	5.0

## 7. Free Water

7.1 *Significance and Use*—The free water analysis determines the amount of free water contained in the sample as opposed to chemically combined water, and prepares the sample for further analysis.

### 7.2 Procedure:

7.2.1 Weigh a sample of the material as received of not less than 50 g to a precision of 0.001 g and spread it out in a thin layer in a suitable vessel. Place in an oven and dry at  $45 \pm 3^\circ\text{C}$  for 2 h, then cool in a desiccator and weigh again. The loss of weight corresponds to the free water.

7.2.2 Retain the sample in a sealed container or in the desiccator for further analysis.

7.3 *Calculation and Report*—Calculate and report loss in weight as a percentage of the sample as received or of the dried sample as required.

7.4 *Precision and Bias*—Neither the precision nor the bias for the free water analysis has been determined.

## 8. Combined Water

8.1 *Significance and Use*—The combined water analysis determines the percent of chemically combined water and is used to calculate the purity of gypsum or the amount of gypsum or gypsum plaster in gypsum products.

8.2 *Interferences*—Some organic materials may partially decompose and give high results. If hydrated compounds other than gypsum are present they may decompose and give high results. Take care that the oven used does not exceed the maximum temperature required, or some carbonates, if present, may decompose and give high results.

### 8.3 Procedure:

8.3.1 For each sample, place three weighing bottles with lids in the preheated calcining oven or furnace and heat for 2 h at 215 to 230°C. Place in the desiccator and allow to cool to room temperature. Weigh the bottles and lids to the nearest 0.0001 g and record the tare weights.

8.3.2 Weigh out three specimens of approximately 1 g each of the sample as prepared in Section 4 and dried in Section 7 to a precision of 0.0001 g in the previously tared weighing bottles and record the total weight with lids.

8.3.3 Place the specimens in the calcining oven with the lids placed loosely on each bottle or crucible for 2 h or until constant weight has been obtained.

8.3.4 Place the lids tightly on the weighing bottles, remove from the oven, and place in the desiccator to cool to room temperature.

8.3.5 Weigh each specimen to a precision of 0.0001 g and record the weights.

8.3.6 Retain the residues for carbon dioxide analysis.

8.4 *Calculation and Report*—Calculate and report the average loss in weight of the three specimens as a percentage of the sample as received or of the dried sample, as required, to the nearest 0.001 g and record the tare weights.

8.5 *Precision and Bias*—Neither the precision nor the bias for the combined water analysis has been determined.

## 9. Carbon Dioxide

9.1 *Summary of Test Method*—The sample is decomposed with HCl and the liberated  $\text{CO}_2$  is passed through a series of scrubbers to remove water and sulfides. The  $\text{CO}_2$  is absorbed with Ascarite, a special sodium hydroxide absorbent,<sup>9</sup> and the gain in weight of the absorption tube is determined and calculated as percent  $\text{CO}_2$ .

9.2 *Significance and Use*—The carbon dioxide analysis is useful in estimating carbonates and organic carbon for chemical balance.

### 9.3 Special Reagents:

9.3.1 *Magnesium Perchlorate Desiccant*<sup>10</sup>—for drying.

9.3.2 *Sodium Hydroxide Absorbent*<sup>9</sup>—a coarse sodium hydroxide coated silica.

9.4 *Special Apparatus*—The apparatus illustrated in Fig. 1 consists of the following:

9.4.1 *Purifying Jar A*, Fleming, containing sulfuric acid.

9.4.2 *Drying Tube B*, U-shaped with side arms and glass-stoppers. Side arms are shaped to hold rubber tubing. Contains Anhydron on left side and Ascarite on right side.

9.4.3 *Erlenmeyer Flask C*, 250 mL, 24/20 ground-glass joint.

9.4.4 *Separatory Funnel D*, with ground glass stopper and interchangeable hollow ground-glass joint. A delivery tube bent at the end extends into the sample flask approximately 15 mm from the bottom and is used to introduce acid into flask.

9.4.5 *Condenser E*.

9.4.6 *Gas-Washing Bottle F*, 250 mL, with fritted disk containing distilled water to retain most of the acid volatilized from the alkalimeter.

9.4.7 *U-Tube G*, containing mossy zinc to remove the last traces of HCl.

9.4.8 *Gas Washing Bottle H*, 250 mL, with fritted disk, containing concentrated  $\text{H}_2\text{SO}_4$  and trap *I*, to remove any  $\text{SO}_3$

<sup>9</sup> Ascarite, manufactured by Arthur H. Thomas has been found satisfactory for this purpose.

<sup>10</sup> Anhydron, manufactured by J. T. Baker Inc. has been found satisfactory for this purpose.

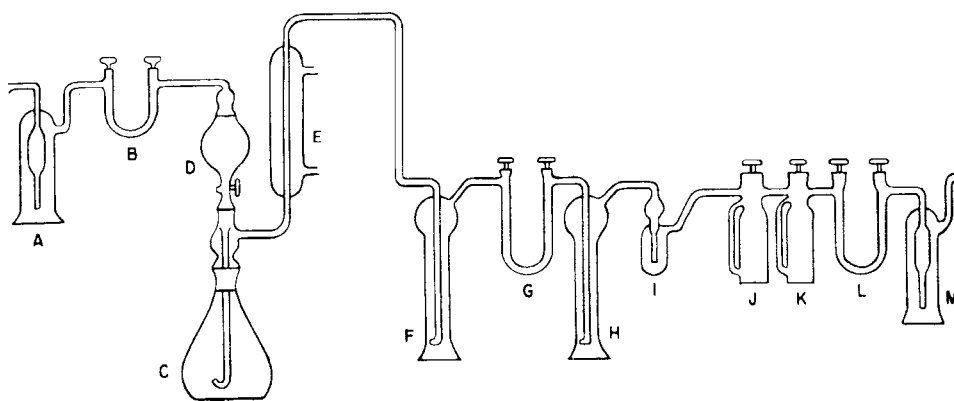


FIG. 1 Apparatus for Carbon Dioxide Analysis

mist that is carried over.

9.4.9 *Absorption Bulb J*, containing Anhydron to remove last traces of water vapor.

9.4.10 *CO<sub>2</sub> Absorption Bulb*, containing Ascarite filled as follows: On the bottom of the bulb, place a layer of glass wool extending above the bottom outlet and on top of this a layer of Anhydron approximately 10 mm thick; immediately above this place another layer of glass wool, then add Ascarite to almost fill the bulb. Place a top layer of Anhydron approximately 10 mm thick on top of the Ascarite and top it off with a covering of glass wool.

9.4.11 *U-Guard Tube L*, filled with Anhydron in left side and Ascarite in right side.

9.4.12 *Purifying Jar M*, Fleming, containing H<sub>2</sub>SO<sub>4</sub>.

9.5 *Procedure:*

9.5.1 After drying as described in Section 8 place the residue obtained in the 250 mL Erlenmeyer flask (C). Connect the flask to the apparatus as shown in Fig. 1. Purge the system free of carbon dioxide by passing a current of CO<sub>2</sub>-free air through the apparatus for 10 to 15 min.

9.5.2 Weigh the absorption bulb to 0.0001 g and attach it to the train. Remove the glass stopper from the separatory funnel, place 50 mL of dilute HCl (1 + 1) in the separatory funnel (D) and replace the stopper with the interchangeable hollow ground-glass joint through which passes a tube for admitting purified air. Open the stopcock of the separatory funnel and admit air through the top of the funnel to force the hydrochloric acid into the Erlenmeyer flask (C).

9.5.3 Start cold water circulating through the condenser (E) and, with CO<sub>2</sub>-free air passing at a moderate rate through the absorption train, place a small hot plate or gas burner under the sample flask and boil for approximately 2 min. Remove the hot plate and continue the flow of purified air at approximately three bubbles per second for 10 min to sweep the apparatus free of CO<sub>2</sub>. Close the absorption bulb, disconnect it from the train and weigh, opening the stopper momentarily to equalize the pressure. Use a second absorption bulb as counterpoise in all weighings unless a single pan balance is used.

9.6 *Calculation*—Calculate the percent CO<sub>2</sub> to the dried sample as follows:

$$\text{Percent CO}_2 = ((A - B)/C \times 100)(1 - D) \quad (1)$$

where:

- A = mass of absorption bulb + CO<sub>2</sub> g,
- B = mass of absorption bulb before the run, g,
- C = mass of specimen, g, and
- D = percent combined water as determined in Section 8 as a decimal.

Calculate the percent CO<sub>2</sub> to the sample as received as follows:

$$\text{Percent CO}_2 = E(1 - F) \quad (2)$$

where:

- E = result of Eq 1, and
- F = percent free water as determined in Section 7 as a decimal.

9.7 *Precision and Bias*—Neither the precision nor the bias for the carbon dioxide analysis has been determined.

## 10. Silicon Dioxide and Other Acid Insoluble Matter

10.1 *Summary of Test Method*—The gypsum and other acid soluble components of the sample are dissolved in dilute hydrochloric acid (HCl). The residue is weighed and calculated as silicon dioxide and other acid insoluble matter.

10.2 *Significance and Use*—The silicon dioxide and other acid insoluble matter analysis determines and is used to report the percentage of one of the inert impurities in gypsum and gypsum products.

10.3 *Procedure*—Perform in triplicate.

10.3.1 Weigh approximately 1 g of the specimen prepared in Section 4 to the nearest 0.0001 g.

10.3.2 Place the specimen in a porcelain casserole. Add approximately 50 mL of 1 + 5 hydrochloric acid. Evaporate slowly and carefully to apparent dryness on a hot plate. The evaporation should take approximately 20 min. Make a blank determination with one casserole in parallel. Cool to room temperature.

10.3.3 Add enough hydrochloric acid (sp gr 1.19) to wet the solid residue. Add 20 mL of water, boil and filter through filter paper.<sup>7</sup> Wash the filter paper thoroughly using not less than 100 mL of room temperature water to render the precipitate chloride free. The most effective washing technique is to use many small quantities of wash water rather than fill the funnel to the brim two or three times. Test the filtrate for chloride by collecting a small amount and adding a few drops of 0.1 normal silver nitrate (AgNO<sub>3</sub>) solution. A white precipitate



indicates more washing is needed. Discard this test solution.

10.3.4 Place all the filtrate back in the same casserole. Evaporate to dryness and heat to 120°C for 1 h and cool. To the cooled casserole add enough HCl (sp gr 1.19) to wet the solid residue. Add 50 mL of water and boil.

10.3.5 Wash the second contents of the casserole through another filter<sup>7</sup> paper. Thoroughly wash the residue in the filter paper until chloride free as in 10.3.3. Retain the filtrate for the iron and aluminum oxide analysis.

10.3.6 Dry sufficient crucibles by placing in a cold muffle furnace during warm up or by placing in a drying oven for 15 to 20 min, then placing in a 900°C muffle furnace. Cool crucibles to room temperature in a desiccator.

10.3.7 Transfer both filter papers to a tared crucible and char slowly without flaming. Burn off all the carbon and ignite in a muffle furnace at 900°C for 15 min.

10.3.8 Cool the crucibles in a desiccator and weigh to the nearest 0.0001 g.

10.4 *Calculation and Report*—Calculate the average weight of the three precipitates and report as silicon dioxide (SiO<sub>2</sub>) and other insoluble matter to the percentage of sample as received or to the dried sample as required.

10.5 *Precision and Bias*—Neither the precision nor the bias for the silicon dioxide and other acid insoluble matter has been determined.

## 11. Iron and Aluminum Oxides

11.1 *Significance and Use*—The iron and aluminum oxides (Fe<sub>2</sub>O<sub>3</sub> + Al<sub>2</sub>O<sub>3</sub>) analysis is used to determine the quantity of these metal oxides in gypsum or gypsum products.

11.2 *Procedure*—To the filtrate, obtained as described in Section 10, add a few drops of nitric acid (HNO<sub>3</sub>), and boil to ensure oxidation of the iron. Add 2 g of ammonium chloride (NH<sub>4</sub>Cl) previously dissolved in water. Make alkaline with ammonium hydroxide (NH<sub>4</sub>OH). Digest hot for a few minutes until the precipitate coagulates. Filter, wash, ignite the precipitate at 1000°C for 30 min or to constant weight in a muffle furnace and weigh as Fe<sub>2</sub>O<sub>3</sub> + Al<sub>2</sub>O<sub>3</sub>. Save the filtrate for the CaO analysis.

NOTE 2—The addition of a pinch of ashless filter paper pulp will aid in the filtration of the precipitate.

11.3 *Calculation*—Calculate Fe<sub>2</sub>O<sub>3</sub> + Al<sub>2</sub>O<sub>3</sub> to the percentage of sample as received or the dried sample as required. This precipitate may be further treated to separate the two oxides, but this is generally unnecessary.

11.4 *Precision and Bias*—Neither the precision nor the bias for the iron and aluminum oxides analysis has been determined.

## 12. Calcium Oxide

12.1 *Significance and Use*—The calcium oxide (CaO) analysis is used to determine the amount of CaO and calculate the amount of calcium carbonate (CaCO<sub>3</sub>) in gypsum and gypsum products.

12.2 *Procedure*:

12.2.1 To the filtrate obtained as described in Section 11 add 5 g of ammonium oxalate ((NH<sub>4</sub>)<sub>2</sub>C<sub>2</sub>O<sub>4</sub>) dissolved in water. Digest hot for 30 min, making sure that the solution is always

alkaline with NH<sub>4</sub>OH. Filter, wash, and ignite the precipitate at 1000°C for 2 h to constant weight in a platinum crucible in a muffle furnace.

12.2.2 *Alternative Method*—To the filtrate obtained as described in Section 11 add 5 g of (NH<sub>4</sub>)<sub>2</sub>C<sub>2</sub>O<sub>4</sub> dissolved in water. Digest hot for 30 min, making sure that the solution is always alkaline with NH<sub>4</sub>OH. Filter and wash, transfer the precipitate to a beaker, and wash the filter paper with hot H<sub>2</sub>SO<sub>4</sub> (1 + 6), catching the washing in the same beaker. Heat gently to complete solution, adding more H<sub>2</sub>SO<sub>4</sub> if necessary. While still warm, titrate with potassium permanganate (KMnO<sub>4</sub>) solution (5.6339 g/L) until the pink color persists.

12.3 *Calculation*—The number of milliliters of KMnO<sub>4</sub> solution used gives directly the percentage of lime in the dried sample. Calculate the CaO to the percentage of sample as received or the dried sample as required.

12.4 *Precision and Bias*—Neither the precision nor the bias for the calcium oxide analysis has been determined.

## 13. Magnesium Oxide

13.1 *Significance and Use*—The magnesium oxide (MgO) analysis is used to determine the amount of MgO and calculate the amount of magnesium carbonate MgCO<sub>3</sub> in gypsum and gypsum products.

13.2 *Procedure*—To the filtrate obtained as described in 12.2.1 or 12.2.2, add enough water to give a total volume of approximately 600 mL. Cool, and add 10 mL of NH<sub>4</sub>OH and 5 g of sodium ammonium phosphate (NaNH<sub>4</sub>HPO<sub>4</sub>) dissolved in water. Stir vigorously until a precipitate begins to form. Let stand overnight. Filter, using a Gooch crucible, and wash with NH<sub>4</sub>NO<sub>3</sub> solution. Ignite at 1000°C for 2 h to constant weight in a muffle furnace.

13.3 *Calculation*—Multiply this weight by 0.36207 to find the weight of magnesium oxide (MgO). Calculate the MgO to the percentage of sample as received or to the dried sample as required.

13.4 *Precision and Bias*—Neither the precision nor the bias for the magnesium oxide analysis has been determined.

## 14. Sulfur Trioxide

14.1 *Summary of Test Method*—In this test method, sulfate is precipitated from an acid solution of the gypsum with barium chloride (BaCl<sub>2</sub>). The precipitate is filtered and weighed as barium sulfate (BaSO<sub>4</sub>) and the sulfur trioxide (SO<sub>3</sub>) equivalent is calculated.

14.2 *Significance and Use*—The specification for gypsum and some gypsum products specifies the amount of calcium sulfate (CaSO<sub>4</sub>) required, either in the dihydrate (CaSO<sub>4</sub>·2H<sub>2</sub>O) or hemihydrate (CaSO<sub>4</sub>·½ H<sub>2</sub>O) form. This procedure assumes that an insignificant amount of sulfate other than calcium sulfate is present. This test method is used to determine compliance to the gypsum and gypsum product specifications. It is also commonly used in quality control work.

14.3 *Interference*—This test method has been developed for natural gypsum and for impurities generally found associated with natural gypsum. Synthetic gypsum may have an additional number of interfering elements and compounds, consequently, this procedure may not give accurate results and is not

recommended. This test method has a number of interferences that theoretically affect the results. Co-precipitation and occlusion can be problems if the solution is either too acidic or too basic. Calculations using SO<sub>3</sub> analysis are most accurate on samples that are known to be completely hydrated or completely dehydrated.

14.4 Procedure:

14.4.1 Having properly selected and prepared the samples as specified in Section 4, weigh a representative specimen of approximately 0.5 g, to the nearest 0.0001 g.

14.4.2 Place the weighed sample into a 400-mL beaker. Add 50 mL of HCl (1 + 5). Boil and disperse with the flattened end of a glass rod while stirring until the sample is completely broken down. Add approximately 100 mL boiling water and continue boiling for 15 min, with this step to be extended as required, so the combined boiling time is not less than 1 h.

14.4.3 Using filter paper<sup>7</sup>, filter into a clean 600-mL flask and rinse the 400-mL beaker thoroughly with hot distilled water. Carefully wash the sides of the 400-mL beaker while wiping the insides with a rubber-tipped glass rod making sure all splatters and insoluble are washed into the filter paper. Dry and burn off the filter paper leaving the residue that can be dried and weighed for insoluble matter should this test method not be otherwise conducted.

14.4.4 Dilute the filtrate to 400 to 500 mL. Add 1 to 2 drops of 0.1 % methyl red indicator. Prepare a 400 to 500-mL sample of 0.05 to 0.1 N HCl. Add 1 to 2 drops of 0.1 % methyl red indicator. Compare the color of this solution to the color of the filtrate. Dilute the filtrate or add HCl (1 + 5) solution as necessary to match the pH of the 0.05 to 0.1 N HCl solution.

14.4.5 Boil the filtrate solution and add 20 mL of near-boiling 10 % barium chloride solution, preferably with the help of a pipette, drop by drop while stirring. The barium chloride solution should be prepared not less than one day before use. Continue boiling the solution for 10 to 15 min and digest hot for 3 h or until the precipitate settles.

14.4.6 Filter<sup>7</sup> and wash with hot water to render the precipitate chloride free. One hundred twenty five to 150 mL of distilled water should be adequate. The filtrate can be tested for chloride by collecting a small amount and adding a few drops of 0.1 N AgNO<sub>3</sub> solution. A white precipitate indicates more washing is needed. Alternately, filtering crucibles<sup>11</sup> may be used for quick filtering if the particular crucibles to be used are tested prior to use by refiltering the filtrate from the crucibles with filter paper, and no more than 2 mg is collected on the filter paper.

14.4.7 Ignite the precipitate and paper in a tared crucible, and slowly char the paper without inflaming. Burn off all the carbon and ignite in a muffle furnace at 800 to 900°C or using bright red heat over a Bunsen burner for 15 to 20 min. Dry the filtering crucibles by placing in a cold muffle furnace during warm-up or in a drying oven prior to igniting in a muffle furnace at 800 to 900°C for 15 to 20 min.

NOTE 3—Thoroughly cleans crucibles before each use and heat in a furnace at 800 to 900°C and cool in a desiccator before taring.

14.4.8 Cool all crucibles in a desiccator and weigh to the nearest 0.0001 g.

14.5 Calculation—Multiply the weight of the precipitate by 0.343 to determine the weight of sulfur trioxide (SO<sub>3</sub>). Calculate the SO<sub>3</sub> to the percentage of sample as received or to the dried sample as required.

14.6 Precision and Bias—Neither the precision nor the bias for the sulfur trioxide analysis has been determined.

15. Chlorides

15.1 Significance and Use—Small amounts of chloride in gypsum or gypsum products can have a detrimental effect on their use. This procedure is used to measure the amount of chlorides present and report it as sodium chloride.

15.2 Procedure:

15.2.1 Weigh approximately 20.0 g of sample as prepared in Section 4 to 0.001 g and transfer to a 400-mL beaker. Add 150 mL of water, stir, and heat to just below the boiling point. Cover with a watch glass and maintain at just below boiling (not less than 80°C) for 1 h with occasional stirring. Filter with suction on a Buchner funnel fitted with a medium filter paper. Wash the residue with four 20-mL portions of hot water.

15.2.2 Add 2 drops of phenolphthalein indicator solution to the filtrate. If the filtrate fails to turn pink, add 0.1 N NaOH solution dropwise with stirring until a faint pink color develops. Add 0.1 N HNO<sub>3</sub> dropwise until the pink color just disappears.

15.2.3 If the chloride content is very low, transfer the entire filtrate quantitatively to a 400-mL beaker and proceed as described in 15.2.4. If larger amounts of chloride are expected, transfer the filtrate quantitatively to a 250-mL volumetric flask, cool to room temperature, and dilute to 250 mL. Take a suitable aliquot, transfer to a 400-mL beaker, and dilute to a volume of 100 to 250 mL.

15.2.4 Place the beaker containing the sample on a white surface, add 0.5 mL (10 drops) of K<sub>2</sub>CrO<sub>4</sub> solution and titrate with AgNO<sub>3</sub> solution using a micro buret having a 10-mL capacity and graduated in divisions of 0.02 mL. Titrate until a faint but definite orange color is visible.

15.2.5 Perform a blank titration using the same volume of water as the sample volume and the same amount of K<sub>2</sub>CrO<sub>4</sub> solution. Titrate to the same color as obtained with the sample.

15.3 Calculation—Subtract the volume of AgNO<sub>3</sub> solution used for the blank titration from the volume used for the sample to give the net titration. A 1-mL net titration is equivalent to 0.002923 g of sodium chloride (NaCl). Calculate the NaCl as a percentage of the sample as received or the dried sample as required.

15.4 Precision and Bias—Neither the precision nor the bias for the chloride analysis has been determined.

16. Report

16.1 Report the results obtained in the analysis as follows:

	Percent
Free water	...
Combined water	...
Carbon dioxide (CO <sub>2</sub> )	...
Silicon dioxide (SiO <sub>2</sub> ) and insoluble matter	...
Iron and aluminum oxides (Fe <sub>2</sub> O <sub>3</sub> + Al <sub>2</sub> O <sub>3</sub> )	...
Lime (CaO)	...

<sup>11</sup> Gooch or Coors filtering crucibles have been found suitable for this purpose.

Magnesium oxide (MgO)	...
Sulfur trioxide (SO <sub>3</sub> )	...
Sodium chloride (NaCl)	...
Total	100.00±

NOTE 4—Since it is frequently advisable to recalculate the results obtained in the chemical analysis in order that they may be more enlightening, the following is submitted for consideration:

(1) Multiply the percentage of combined water by 4.778 to obtain purity or percentage gypsum. To calculate the percentage of CaSO<sub>4</sub>·½ H<sub>2</sub>O in plasters, multiply the percentage of gypsum by 0.8430.

(2) Multiply the percentage of combined water by 2.222 to obtain the amount of SO<sub>3</sub> combined as gypsum.

(3) Subtract the result obtained in (2) from the total SO<sub>3</sub> found by analysis to obtain the excess SO<sub>3</sub>.

(4) Multiply the excess SO<sub>3</sub> by 1.700 to obtain the percentage anhydrite, CaSO<sub>4</sub>.

(5) Multiply the percentage of gypsum found in (1) by 0.3257 to obtain the percentage of CaO combined as gypsum.

(6) Multiply the percentage of anhydrite found in (4) by 0.4119 to obtain the percentage of CaO combined as anhydrite.

(7) Add (5) and (6) together. Then subtract this result from the total CaO percentage found by analysis.

(8) Multiply the excess CaO percentage by 1.785 to obtain the percentage of calcium carbonate.

(9) Multiply the percentage of MgO by 2.091 to obtain the percentage of magnesium carbonate.

NOTE 5—Having made the calculations in Note 4, the results may be reported as follows:

	Percent
Gypsum (CaSO <sub>4</sub> ·2H <sub>2</sub> O)	...
Anhydrite (CaSO <sub>4</sub> natural and manufactured) (Note 4)	...
Silicon dioxide and insoluble (SiO <sub>2</sub> + Ins.)	...
Iron and aluminum oxide (R <sub>2</sub> O <sub>3</sub> )	...
Calcium carbonate (CaCO <sub>3</sub> )	...
Magnesium carbonate (MgCO <sub>3</sub> )	...
Sodium chloride (NaCl)	...
Total	100.00±

NOTE 6—The presence of the different forms of CaSO<sub>4</sub> may be determined by a microscopic examination. A paper titled "Gypsum Analysis with the Polarizing Microscope" containing suggested methods can be found in ASTM STP 861.<sup>12</sup>

## ALTERNATIVE PROCEDURE FOR ANALYSIS FOR CALCIUM SULFATE BY THE AMMONIUM ACETATE METHOD<sup>13</sup>

### 17. Significance and Use

17.1 This test method covers the determination of calcium sulfate in gypsum and gypsum products by extraction with ammonium acetate solution, and may be used as an alternative method.

### 18. Reagents and Materials

18.1 *Ammonium Acetate Solution*—Dissolve 454 g of ammonium acetate in 2 L of water. Add sufficient NH<sub>4</sub>OH to make the solution distinctly ammoniacal, using phenolphthalein as the indicator.

18.2 *Ammonium Hydroxide Wash Solution*—Dilute 100 mL of concentrated ammonium hydroxide (NH<sub>4</sub>OH, sp gr 0.90) to 1 L with water.

18.3 *Filter Aid*—Diatomaceous silica, analytical grade.

18.4 *Phenolphthalein Indicator Solution*.

### 19. Procedure Using Gooch Crucible

19.1 Weigh rapidly approximately 4 g of the well-mixed sample and transfer to a 600-mL beaker. Make all weighings to 0.001 g, except weigh the crucibles and their contents to 0.0001 g.

19.2 Without delay, weigh approximately 1 g of the well-mixed sample in a tared weighing bottle having a ground-glass stopper. Dry the sample and weighing bottle to constant weight at 45°C. Stopper weighing bottles immediately upon removal from the oven in order to prevent absorption of moisture from the air upon cooling.

19.3 If the percentage by weight of combined water held by the calcium sulfate is desired, heat the sample and weighing bottle to constant weight at 220°C.

19.4 To the contents of the 600-mL beaker (19.1), add 350 mL of the ammonium acetate solution, and stir the mixture thoroughly to loosen all of the solid matter from the bottom of the beaker. Add 0.2000 g of redried diatomaceous silica to the mixture, heat the beaker and contents to 70°C on a steam bath, and maintain at that temperature for 30 min, while stirring frequently. During heating, keep the solvent ammoniacal by additions of NH<sub>4</sub>OH and phenolphthalein, if indicated. Meanwhile, heat a supply of the ammonium acetate solution to 70°C, keeping it also distinctly ammoniacal. Filter the mixture, with suction, through a tared Gooch crucible, stirring frequently during filtration to keep the diatomaceous earth suspended in the liquid. Wash the Gooch crucible containing the residue with five 10-mL portions of the warm acetate solution, draining thoroughly after each washing. Wash in the same manner with eight 10-mL portions of the NH<sub>4</sub>OH wash solution. Take care to wash the upper walls of the Gooch crucible. Drain the crucible dry with suction, place in an oven at 70°C, and dry to constant weight (Note 7). Allow the crucible to cool in a desiccator before weighing.

NOTE 7—Avoid overheating in all oven drying of ammonium acetate residues; that is, place crucibles well away from the heating elements. This is of particular importance for samples high in impurities, as these impurities often have water of hydration that is lost on local overheating.

### 20. Procedure Using Tared Filter Papers

NOTE 8—This procedure is suggested where several samples are to be analyzed at once. It has been found that gravity filtration on six samples will proceed as rapidly as it is possible to handle the samples.

20.1 Dry a quarter-folded, 110-mm quantitative filter paper overnight at 70°C in a wideform, glass-stoppered, 30 by 60-mm weighing bottle. After drying, cool the weighing bottle and paper in a desiccator, and weigh.

20.2 Treat the sample exactly as described in 19.1, 19.2, and 19.4 prior to the filtration. Filter the mixture by gravity through a 70-mm glass funnel, stirring frequently during filtration to keep the diatomaceous silica suspended in the liquid. Wash the filter paper and residue with five 10-mL portions of warm acetate solution, draining thoroughly after each washing. Wash in the same manner with eight 10-mL portions of the NH<sub>4</sub>OH wash solution. After final draining, replace the paper and residue in the weighing bottle, and dry at 70°C to constant

<sup>12</sup> Green, George W., "Gypsum Analysis with the Polarizing Microscope," *The Chemistry and Technology of Gypsum*, ASTM STP 861, ASTM, 1984, pp. 22–47.

<sup>13</sup> This procedure was developed by L. S. Wells and W. F. Clarke, National Bureau of Standards, and modified by B. E. Kester, United States Gypsum Co.



weight. Cool the weighing bottle, paper, and residue in a freshly prepared desiccator before weighing; this is essential, due to the hygroscopic character of paper.

## 21. Calculation

21.1 Calculate the percentage of loss in weight at 45°C (free water) as follows:

$$\text{Loss in weight at } 45^{\circ}\text{C, \%} = [(A - B)/C] \times 100 \quad (3)$$

where:

- A* = original weight of sample and weighing bottle,
- B* = weight of sample and weighing bottle dried to constant weight at 45°C, and
- C* = original weight of sample.

Calculate the weight of the 4-g sample (19.1), corrected for loss on heating to constant weight at 45°C.

21.2 Calculate the percentage of combined water as follows:

$$\text{Combined water, \%} = [(B - D)/(B - E)] \times 100 \quad (4)$$

where:

- B* = weight of sample and weighing bottle dried to constant weight at 45°C,
- D* = weight of sample and weighing bottle dried to constant weight at 220°C, and
- E* = weight of weighing bottle.

21.3 Calculate the percentage of  $\text{CaSO}_4 \cdot \text{XH}_2\text{O}$  on the basis of the sample dried to constant weight at 45°C as follows:

$$\text{CaSO}_4 \cdot \text{XH}_2\text{O, \%} = [F - (G - H)/F] \times 100 \quad (5)$$

where:

- F* = weight of sample, corrected for loss on heating to constant weight at 45°C,
- G* = weight of dried crucible and contents (19.4) or weight of weighing bottle and contents (20.2), and
- H* = weight of crucible plus diatomaceous silica used as filter aid (19.4), or weight of weighing bottle, diatomaceous silica used as a filter aid and the weight of filter paper (20.2).

## 22. Precision and Bias

22.1 Neither the precision nor the bias for the analysis of calcium sulfate by the ammonium acetate method has been determined.

### ALTERNATIVE PROCEDURE FOR ANALYSIS FOR SODIUM CHLORIDE BY THE COULOMETRIC METHOD<sup>14</sup>

## 23. Significance and Use

23.1 This test method covers the determination of sodium chloride in gypsum and gypsum products by the coulometric method, and may be used as an alternative method.

## 24. Interferences

24.1 The presence of sulfide, sulfhydryl, or other silver reactive substances will lead to high results. Such interfering substances may be removed by alkaline oxidation with hydrogen peroxide.

## 25. Apparatus

### 25.1 Chloride Meter:

25.1.1 The instrument shall be equipped to measure the concentration of dissolved chloride in aqueous solutions by the coulometric method.<sup>15</sup>

25.1.2 The instrument shall be capable of measuring chloride concentrations in the range from 10 to 260 mg/L with a repeatability of  $\pm 1$  mg/L.

## 26. Reagents

26.1 *Acid Buffer Solution*—Dissolve 100 mL of 99.5 % acetic acid ( $\text{HC}_2\text{H}_3\text{O}_2$ ) and 5.5 mL of concentrated nitric acid (sp gr 1.42) in approximately 200 mL of water and dilute to 500 mL.

26.2 *Diluted Standard Solution* (100 mg Cl/L)—Dilute 5.00 mL of stock standard solution to 500 mL.

26.3 *Gelatin Solution*—Add 2.5 g of gelatin and 0.5 g of thymol blue to 250 mL of water and dissolve by stirring continuously while bringing to a boil. With the solution just boiling, continue stirring until all the thymol blue is dissolved. Add 0.5 g of thymol, cool, and dilute the solution to 500 mL.

NOTE 9—The gelatin solution holds the precipitated silver chloride ( $\text{AgCl}$ ) in suspension and also indicates the presence of the acid buffer. The solution will keep for 3 months at room temperature or longer if refrigerated. Warm the refrigerated solution to room temperature before use.

26.4 *Stock Standard Solution* (10 g Cl/L)—Dissolve 8.240 g of dried sodium chloride ( $\text{NaCl}$ ) in water and dilute to 500 mL.

## 27. Procedure

27.1 Weigh 20.0 g of the well-mixed sample and transfer to a 150-mL beaker.

27.2 Add 50 mL of water, boil, allow the solid material to settle, and filter off the solution. Add an additional 50 mL of water to the solids, boil, and pour the contents of the beaker into the filter. Wash the residue with 100 mL of hot water, adding the washing to the filtrate. Cool and dilute with water to 250 mL.

27.3 Switch on the chloride meter and allow a period of 25 min before use. Set the counter to zero.

27.4 Place a magnetic stirring bar in the test beaker, add 10 mL of diluted standard solution, 3 mL of acid buffer solution, and 5 drops of gelatin solution. Place the test beaker on the platform and lower the electrodes into the solution. Press the “start” button until the pilot light is extinguished. The counter will begin to register after a few seconds. Do not remove the electrodes from the sample until the pilot light comes on. Read the chloride content from the counter. The reading should be  $100 \pm 1$  mg Cl/L. If this reading is not obtained, refer to the manufacturer’s instruction manual. Reset the counter to zero.

27.5 Repeat the procedure used in 27.4, using 10 mL of the sample solution instead of the diluted standard solution. Read

<sup>15</sup> The EEL Chloride Meter, available from Fisher Laboratory Supplies Co., has been found satisfactory. Other instruments available for the coulometric determination of chloride are the Aminco Chloride Titrator, available from American Instrument Co. Ltd.; the Buchler Chloridometer, available from Buchler Instruments Division of Nuclear Chicago Corp.; and the Fiske/Marius Chlor-O-Counter available from Johns Scientific.

<sup>14</sup> This procedure was developed by WestroC Industries Limited.



the result as milligrams of chlorine per litre. When all tests are completed, lower the electrodes into reagent water.

## 28. Calculation

28.1 Calculate the amount of NaCl as a percentage of the sample as received or dried sample as follows:

$$\text{NaCl, \%} = 0.00206 \times A \quad (6)$$

where:

A = chloride meter reading, mg Cl/L.

## 29. Precision and Bias

29.1 Neither the precision nor the bias for the sodium chloride analysis by the coulometric method has been determined.

### DETERMINATION OF SAND IN SET PLASTER

## 30. Summary of Test Method

30.1 This test method for the determination of the sand content of set gypsum plaster requires for accurate results the following determinations:

30.1.1 Determination of the percentage of insoluble matter in the sand used with the plaster,

30.1.2 Determination of the percentage of insoluble matter in the gypsum neat plaster, and

30.1.3 Determination of the percentage of insoluble matter in the sanded calcined plaster.

NOTE 10—If samples of the original gypsum neat plaster and the sand are not available, an approximation of the insoluble matter may be obtained by the use of this method on plaster and sand from the same sources as those from which the plaster to be analyzed was originally prepared.

## 31. Significance and Use

31.1 This test method is used for determining the sand content of samples of aggregated plaster taken from job sites to determine compliance with Specification C 842.

## 32. Reagents

32.1 *Ammonium Acetate* (250 g/L)—Dissolve 250 g of ammonium acetate ( $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ ) in water and dilute to 1 L.

32.2 *Ammonium Hydroxide* (1 + 59)—Mix 1 volume of concentrated ammonium hydroxide ( $\text{NH}_4\text{OH}$ ) (sp gr 0.90) with 59 volumes of water.

## 33. Sampling

33.1 Where plaster to be tested is part of a two-coat or three-coat plastering operation, take the sample for analysis from that portion of the entire plaster sheet that comprises the single coat being tested. Separate succeeding coats of plaster by use of a stiff putty knife or similar implement. Not less than 500 g shall be taken as a sample, the sample preferably being obtained from different sections of the wall or ceiling under examination.

## 34. Procedure

34.1 In a clean porcelain mortar, grind the set plaster sample to the size of the largest sand particles present, or smaller, so

that approximately 100 % of the sample will pass a 2.36-mm (No. 8) sieve. Fine grinding makes solution of the gypsum faster. Place approximately 200 g of the ground sample in a porcelain casserole or evaporating dish, and calcine on a sand bath. Stir the sample continuously with a thermometer during the heating, and adjust the rate of heating so that 20 to 30 min will be required to raise the temperature of the sample to  $160 \pm 5^\circ\text{C}$ . Cool the sample to room temperature in a desiccator.

34.2 After cooling, weigh accurately  $20 \pm 0.05$  g of the calcined sample into a 600-mL beaker. Add 300 to 350 mL of  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  solution, which should be slightly alkaline to litmus paper. If acidic, add a few millilitres of  $\text{NH}_4\text{OH}$  (1 + 59) to the stock  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  solution to render it slightly alkaline prior to the addition to the test sample.

34.3 Warm the suspension to a temperature of  $70 \pm 5^\circ\text{C}$  and stir continuously for 20 to 30 min. Filter the warm suspension with the aid of suction through a small Büchner funnel or Gooch crucible in which filter paper has previously been placed. Refilter the first 100 mL of the filtrate. Wash the sand remaining in the beaker onto the filter with an additional 100 mL of warm ammonium acetate solution. Wash the beaker and residue with 200 to 300 mL of water, dry the funnel and sand at  $100^\circ\text{C}$  to constant weight. The weight of the residue is the weight of insoluble matter.

34.4 *Insoluble Matter in Sand*—Determine the weight of insoluble matter in sand as described in 34.1-34.3, except that no grinding of the sample is necessary.

34.5 *Insoluble Matter in Gypsum Neat Plaster*—Determine the weight of insoluble matter in the gypsum neat plaster as described in 34.1-34.3, except that only a 5-g sample is required and no grinding of the sample is necessary.

## 35. Calculation

35.1 *Insoluble Matter in Plaster*—Multiply by 5 the weight of the insoluble matter obtained as described in 34.3 to obtain the percentage of insoluble matter in sanded plaster.

35.2 *Insoluble Matter in Sand*—Multiply by 5 the weight of the insoluble matter in sand obtained as described in 34.4 to obtain the percentage of insoluble matter in sand.

35.3 *Insoluble Matter in Gypsum Neat Plaster*—Multiply by 20 the weight of the insoluble matter in gypsum neat plaster obtained as described in 34.5 to obtain the percentage of insoluble matter in gypsum neat plaster.

35.4 Calculate the percentage of sand in the sanded plaster as follows:

$$X = [(C - B)/(A - B)] \times 100 \quad (7)$$

where:

X = % of sand in sanded plaster,

A = % of insoluble matter in the sand,

B = % of insoluble matter in the gypsum neat plaster, and

C = % of insoluble matter in the sanded plaster.

35.5 To express the results as a ratio of the parts of sand per part of plaster by weight, the following equation may be used:

$$\text{Ratio of sand to plaster} = X/(100 - X) \quad (8)$$

NOTE 11—The results obtained by the above procedure indicate the amount of sand originally mixed with the gypsum neat plaster before it had been gaged with water or set.

### 36. Precision and Bias

36.1 Neither the precision nor the bias for the analysis of sand in set plaster has been determined.

### WOOD-FIBER CONTENT IN WOOD-FIBER GYPSUM PLASTER

#### 37. Significance and Use

37.1 This test method is used to determine the weight of wood fiber in wood-fibered plaster.

#### 38. Procedure

38.1 Place a 100-g sample of wood-fiber plaster, prepared as described in Section 4 on a 600- $\mu\text{m}$  (No. 30) sieve<sup>6</sup> nested over a 150- $\mu\text{m}$  (No. 100) sieve.<sup>6</sup> Wash the plaster on the 600- $\mu\text{m}$  sieve with a stream of cold water, removing the 600- $\mu\text{m}$  sieve when the fiber on it is practically or entirely free of plaster. Next, wash the material on the 150- $\mu\text{m}$  sieve until the bulk of the plaster has been washed through the sieve and the residue is mainly fiber. Transfer the material retained on the 150- $\mu\text{m}$  sieve to a 300-mL, vitreous enamel, lipped pan, adding the charge on the 600- $\mu\text{m}$  sieve if the fiber contains any adhering particles of plaster. Elutriate the material in the pan (purify by washing and straining, effecting as clean a separation of fiber from plaster as is feasible), catching the elutriated fibers on a 150- $\mu\text{m}$  sieve. To avoid loss of the fine particles of fiber, it may be necessary to make the transfer from the pan to the 150- $\mu\text{m}$  sieve by several stages of washing, stirring the charge, and quickly pouring upon the sieve the fiber flotations, repeating the elutriation procedure several times. Examine the fiber collected on the 150- $\mu\text{m}$  sieve and repeat the elutriation if it seems desirable.

38.2 Dry the sieves (or sieve, as the case may be) and the residue contained therein overnight in an oven maintained at a temperature of 45°C. Carefully invert the sieves, or sieve, over a piece of white paper, and transfer the residual material to the paper by brushing the bottom of the inverted sieve. Examine the transferred material visually, noting whether the separation of fibers from plaster has been complete. Then transfer the material to a weighed platinum crucible and dry to constant weight at a temperature of 45°C. If the previous visual examination of the charge on the white paper showed that the fiber was practically free of particles of plaster, report as the percentage of fiber the weight of the fiber dried at 45°C, divided by 100. If, on the other hand, the visual examination revealed the presence of an appreciable quantity of plaster associated with the fiber, carefully ignite the contents of the crucible to constant weight. In this case, report as the percentage of fiber the loss on ignition, divided by 100.

#### 39. Precision and Bias

39.1 Neither the precision nor the bias for the analysis of wood-fiber content in wood-fiber gypsum plaster has been determined.

### OPTIONAL PROCEDURE FOR ANALYSIS FOR SODIUM BY THE ATOMIC ABSORPTION METHOD

#### 40. Significance and Use

40.1 This test method covers the determination of sodium in

gypsum and gypsum products by the atomic absorption method, and may be used as an optional procedure.

#### 41. Interferences

41.1 Sodium is partially ionized in the air-acetylene flame. The effects of ionization may be significantly overcome by the addition of 1 to 2 g/L of another alkali to blanks, standards, and samples. Alternatively, the air-hydrogen flame may be used, as it produces less ionization and less visible emission than the air-acetylene flame.

#### 42. Apparatus

##### 42.1 Atomic Absorption Spectrophotometer:

42.1.1 The instrument shall be equipped to measure the concentration of dissolved sodium in aqueous solutions using either the air-acetylene or air-hydrogen flame.

42.1.2 The instrument shall be capable of measuring sodium concentrations within the optimum analytical range of 0.1 to 0.5 absorbance units while providing a coefficient of variation of approximately 0.5 to 2 %.

#### 43. Reagents

43.1 *Solvent*—Use deionized water to prepare all solutions. If an alkali is to be included for the purpose of suppressing sodium ionization, it is most convenient to add it to the solvent at the start. In this way a constant concentration of alkali in blank, standards, and sample solution is ensured.

43.2 *Stock Standard Solution* (1.000 g Na/L)—Dissolve 2.5418 g of dried sodium chloride (NaCl) in water and dilute to 1 L with water.

43.3 *Dilute Standard Solutions*—Prepare dilute standard solutions bracketing the absorbance range of the dilute sample solution, using the stock standard solution. (Solutions having a concentration less than approximately 0.500 g/L are unstable for periods of more than one day.)

#### 44. Procedure

44.1 Take 18 g of the well-mixed sample and transfer to a 150-mL beaker.

44.2 Add 50 mL of water, boil, allow the solids to settle, and decant the supernatant liquid into a filter. Add an additional 50 mL of water to the solids, boil, and pour the contents of the beaker into the filter. Wash the residue with 100 mL of hot water, adding the washing to the filtrate. Cool the filtrate to room temperature and dilute to 500 mL in a volumetric flask to make the stock sample solution. Take 10 mL of the solution and make up to 500 mL in a second volumetric flask, to make the dilute sample solution.

44.3 Determine the absorbance readings on the dilute standard solutions and the solvent blank at a wavelength of 589.0 to 589.6 nm, following the manufacturer's instruction manual. Subtract the absorbance value for the blank from the absorbance values for the dilute standard solutions and prepare a curve relating sodium concentration in milligrams per litre to absorbance values.

NOTE 12—If the absorbance of the dilute sample solution is known to lie within the linear range, that is, the sodium concentration is below approximately 1 mg/L only one standard and the solvent blank are needed to prepare the curve.

44.4 Determine absorbance readings on the dilute sample solution using the same technique, followed with the dilute standard solutions. The concentration of sodium in the dilute sample solution in milligrams per litre is then found by consulting the standard curve.

#### 45. Calculation

45.1 Calculate the amount of NaCl as a percentage of the sample received or the dried sample as required as follows:

$$\text{Sodium calculated as \% NaCl} = A \times 6.3553/S \quad (9)$$

where:

$A$  = concentration of dilute sample solution, mg/L, and

$S$  = weight of sample, g.

#### 46. Precision and Bias

46.1 Neither the precision nor the bias for the analysis of sodium by the atomic absorption method has been determined.

### OPTIONAL PROCEDURE FOR ANALYSIS FOR SODIUM BY FLAME PHOTOMETRY

#### 47. Significance and Use

47.1 This test method covers the determination of sodium in gypsum and gypsum products by flame photometry, and may be used as an optional procedure. This test method is based on Methods D 1428.

#### 48. Interferences

48.1 Radiation interferences caused by elements other than that being determined are the greatest contributor to error in flame photometry. Some effects are positive and others negative. Of the elements encountered in this analysis, the greatest effect is that of one alkali metal on another. The foreign-element effects cannot be entirely compensated for without employing calibration standards closely duplicating the composition of the sample. However, the effects may be minimized by operating at the lowest practical sodium concentration range or by removal of the interfering elements. For example, aluminum has a depressing effect on alkali-metal emission, which may be of serious consequence. Aluminum should be removed from the extraction liquid prior to flame photometry if its concentration has been found, by preliminary tests, to exceed that of the sodium.

48.2 Self-absorption causes the curve of intensity versus concentration to decrease its slope at higher concentrations, tending to reduce accuracy. Bracketing the unknown by known standard solutions tends to minimize this interference.

#### 49. Apparatus and Materials

49.1 *Flame Photometer*—The instrument shall consist of an atomizer and burner; suitable pressure-regulating devices and gages for fuel and air or oxygen; an optical system, consisting of suitable light-dispersing or filtering devices capable of preventing excessive interference from light of wavelengths other than that being measured; and a photosensitive indicating device.

49.2 *Supply of Fuel and Air or Oxygen*—The supplies of fuel and air or oxygen shall be maintained at pressures somewhat higher than the controlled operating pressure of the instrument.

#### 50. Reagents

50.1 Prepare the following stock solutions from reagents that have been dried to constant weight at 105°C. Store the stock solutions in polyethylene or equally alkali-metal-free containers.

50.2 *Sodium Chloride Solution* (2.5418 g/L)—Dissolve 2.5418 g of sodium chloride (NaCl) in water and dilute to 1 L with water. This stock standard solution contains 1.000 g/L of sodium ion.

50.3 *Dilute Standard Solutions*—Prepare dilute standard solutions from stock standard solution, bracketing the expected sodium concentration range of the diluted sample extraction liquid. For example, if the expected range of the sample extraction liquid is between 0 and 0.010 g/L, prepare eleven equally spaced standards in tenths of the maximum.

#### 51. Calibration of Instrument

51.1 Select the proper photocell; the blue-sensitive phototube having a range from 320 to 620 nm is required for sodium determination. Open the slit width to approximately one fourth of the maximum opening, set the instrument to the maximum sensitivity range, and balance the meter to obtain electrical equilibrium.

51.2 Feed fuel and air or oxygen to the burner and ignite the emitted mixture. Adjust fuel and air or oxygen pressures and follow the procedures for warm-up time prescribed by the instrument manufacturer. Set the scale-reading dial at approximately 95 % of full scale, introduce a solution containing the maximum amount of sodium in the range to be covered, and allow the emitted light to strike the photocell.

51.3 Select the proper filter, if a filter-type instrument is used. For instruments employing spectral dispersing devices, turn the wavelength dial back and forth slowly and carefully in the vicinity of 589 nm. The galvanometer can be made to pass through a maximum deflection. This wavelength setting produces maximum sensitivity, and should not be disturbed during the test.

51.4 Continue to atomize the maximum standard of the range to be covered, and set the scale-reading of the dial at exactly full scale (100 or 1000) (Note 13). Adjust the gain so as to balance the galvanometer needle.

NOTE 13—For instruments equipped with a variable slit, carry out the procedure described above with the slit width between fully closed and one fourth open for the preliminary test. In determining the proper slit width for optimum instrument performance, consideration must be given to the fact that the intensity of the emission line is approximately proportional to the slit width, whereas continuous background intensity increases as the square of the slit width. A decrease in slit width results in decreased illumination of the phototube for a given concentration and is compensated for by increasing the gain of the instrument. The most favorable operating conditions are obtained with the smallest slit width that does not result in instability of the galvanometer needle when it is set to give full-scale reading with the maximum standard in the range to be covered. When altering the slit width, determine the background by atomizing a zero standard, and check for sensitivity setting and instrument stability with the maximum standard in the range to be covered. Lower ranges require wider slit widths. Determine and record the optimum slit width for each range and element to be covered. Use these values in all subsequent tests.

51.5 Determine the emission intensity of all standards.

51.6 Plot emission intensity (scale reading) versus concentration on linear graph paper. For the lower ranges, the curve thus prepared approximates a straight line but may not intersect zero because of background intensity. At higher ranges, the curves show a decrease in slope with increasing concentration. Record on graphs all data in regard to slit width, fuel pressure, and air or oxygen pressure.

## 52. Procedure

52.1 Weigh 25 g of the well-mixed sample and transfer to a 150-mL beaker.

52.2 Add 50 mL of water, boil, allow the solid material to settle, and decant the supernatant liquid into a filter. Add an additional 50 mL of water to the solids, boil and pour the contents of the beaker into the filter. Wash the residue with 75 mL of hot water, adding the washing to the filtrate. Cool and dilute with water to 200 mL in a volumetric flask to make the stock sample solution.

52.3 Take 5 mL of the stock sample solution and make up to 100 mL in a volumetric flask to make the dilute sample solution.

NOTE 14—If the concentration of sodium in the sample is found to be greater than the maximum standard, further dilute sample solution with water to bring the concentration within the range. If the concentration of sodium in the sample is less than one tenth of the value of the maximum standard, prepare a new dilute sample solution from the stock sample solution to bring the concentration within the range.

52.4 Turn the instrument on and feed fuel and air or oxygen to the burner. Ignite the gas mixture. For instruments with an

adjustment slit, set the width to the value determined as outlined in Note 13.

52.5 Place the scale-reading dial at maximum. Atomize the sample and allow its emitted light to strike the photocell. Set the wavelength to 589 nm as described and adjust the gain to balance the galvanometer. Determine the emission intensity of the sample.

52.6 Refer to the standard curve prepared above and read off the concentration of sodium ion in the dilute sample solution in milligrams per litre.

## 53. Calculation

53.1 Calculate the amount of NaCl as follows:

$$\text{Sodium, calculated as NaCl, \%} = A \times 0.04067 \quad (10)$$

where:

$A$  = concentration of sodium, mg/L read off the standard curve.

## 54. Precision and Bias

54.1 Neither the precision nor bias of the analysis for sodium by flame photometry has been determined.

## 55. Keywords

55.1 ammonium acetate method; atomic absorption; chemical analysis; coulometric method; flame photometry; gypsum; gypsum concrete; gypsum board; gypsum products; plaster; sand in set plaster; wood-fiber plaster

## SUMMARY OF CHANGES

This section identifies the location of changes to these test methods that have been incorporated since the last issue. Committee C-11 has highlighted those changes that affect the

technical interpretation or use of these test methods.

(1) Section 6.1.20 was revised.

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